

pH Dependent Isotropic to Nematic Phase Transitions in Graphene Oxide Dispersions Reveal Droplet Liquid Crystalline Phases

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Size fractionation, amplified by the surface charge density of graphene oxide (GO) sheets, broadens the pH dependent isotropic (I) to nematic (N) phase transition in aqueous dispersions of graphene oxide (GO). In this biphasic region, a highly organized droplet nematic phase of uniform size ($20 \pm 2.8 \mu\text{m}$ diameter) with an isotropic interior is observed.

Suspensions of 2D sheet-like materials exhibit transition from a disordered isotropic (I) phase to ordered nematic (N) phase. As noted by Onsager,¹ despite their loss of orientational entropy, the N phase, is stabilized by a gain in excluded volume (or configurational) entropy. Graphene - a single atom thick layer of carbon atoms is receiving immense attention owing to the combination of properties such as large intrinsic mobility of electrons, massive surface area, immense mechanical strength and large thermal conductivity. Although significant progress has been made in the solid-state synthesis of graphene, liquid-phase graphene possess opportunities for large-scale synthesis and novel mesophases that may lead to ordered macroscopic structures such as thin films and fibers with enhanced electrical, mechanical and optical properties.^{2,3} Graphene solvated in chlorosulfonic acid² and stabilized by surfactant⁴ exhibit liquid crystalline behavior, as well as their composites with discotic molecules.⁵ Chemically oxidized graphene sheets or graphene-oxide (GO) suspensions in water, the most accessible precursor to graphene-based materials, not surprisingly demonstrate I-N phase transitions, too.⁶⁻¹⁰ These studies have shown that the nematic phases of graphene exhibit 'brush-like' texture in massively organized

continuous domains that may reach length-scales of millimeters. These length-scales are relevant to achieving macroscopically ordered materials and fluid-phase alignment. However, graphene sheets are flexible and although this flexibility is likely to yield isolated droplet liquid crystalline phases, yet such confined phases have not been reported so far. The ability to form highly ordered N phases from GO arises because GO can be stabilized in aqueous solvent by electrostatic repulsive forces between the particles as demonstrated by large (-20 to -40 mV) zeta potential measurements.¹¹⁻¹³ The charged functional groups are primarily carboxylic groups on the edges and hydroxyls on the basal planes, which can be protonated/deprotonated by change in pH, thus surface charging of these particles depend on the pH. pKa of GO is ~ 4 for the carboxylic groups and ~ 9 for the phenolic groups,¹⁴ however GO suspensions can be stable at lower pH through hydrogen bonding between the oxidized groups and water molecules.¹⁵ In addition, the basal plane contains hydrophobic domains that make GO sheets amphiphilic, and the ionisable groups render this amphiphilicity a dependence on pH and size of the GO sheets with smaller GO sheets being more hydrophilic than larger GO sheets.^{13,16-18} It can thus be expected that the I-N phase transitions are dependent upon factors such as salt concentration, pH and the size of GO sheets. Although the effect of salt concentration on the I-N phase transition of GO has been studied,³ the influence of pH on the phase-behavior has not been detailed in the literature. In this study, we examined the influence of pH on the formation of GO liquid crystals in suspensions with high polydispersity (average size: $415 \text{ nm} \pm 77\%$), through determination of classical I-N phase behavior and the structure of the liquid crystalline phases. GO was prepared from graphite by the modified Hummers's procedure^{19,20} (see S1 in ESI) and dispersed in water to form suspensions at nine different concentrations between 1 to 20 mg/ml and five pH values (1, 2, 6, 9 and 14). pH 2 and 6 were chosen to have more than 99% of the groups protonated and deprotonated since they are ± 2 pH units from the pKa ~ 4 of the carboxylic acids groups on GO.^{12,14} The suspensions were sonicated, when their pH was ~ 2 , for 1 min at ~ 10 W power. The pH was adjusted with concentrated NaOH or HCl solutions. The suspensions were loaded into rectangular capillary tubes (VitroCom) having inner cross section of $0.2 \times 2 \text{ mm}$, significantly larger than the GO sheets or the droplets, and then sealed and left to stand vertically for phase separation. Polarized Light Microscopy by LC-PolScope imaging system^{21,22} ($40\times$ 0.75/0.85NA objective) was used to distinguish between the isotropic and nematic phases and determine the fine-structure of the phases. The main advantage of LC-PolScope compared to traditional polarizing microscopes is that it can measure the retardance or diattenuation along with their dominant orientation simultaneously at all points of the image without having to mechanically rotate the

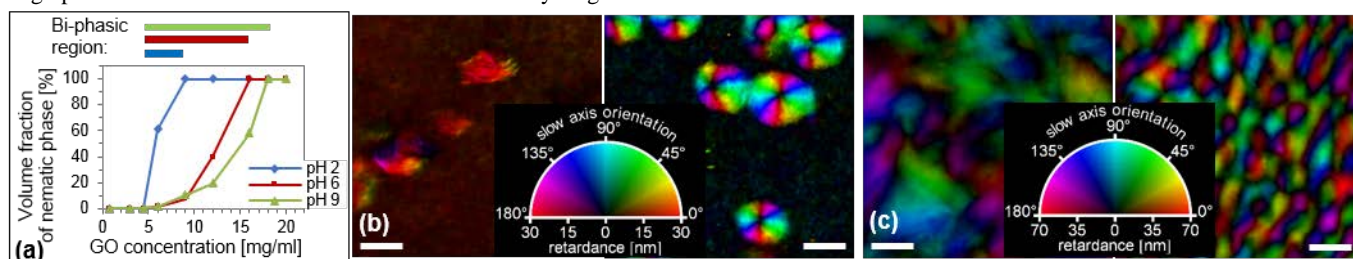


Figure 1. (a) Volume fraction of nematic phase versus concentration at different pH for GO dispersions in water. Pronounced broadening of the I-N biphasic region with increasing pH is noticeable. (b)-(c) LC-PolScope processed images displaying the azimuth and retardance. The color legend is a guide relating hue to azimuth (expressed around the circumference of the semi-circle) and brightness to retardance (expressed radially in nm). Typical morphology of: (b) metastable structures [6 mg/ml, pH 2 (left) and 12 mg/ml, pH 6 (right)], and (c) completely separated nematic phase [at pH 2 (left) and pH 6 (right)]. Scale bars are $20 \mu\text{m}$.

specimen. Note that the birefringence measurement is affected by the diattenuation, but the diattenuation is sensitive only to differential transmittance, as detailed in Mehta *et al.*²³ and demonstrated for reduced-GO in Tkacz *et al.*²⁴ In the two extreme cases of pH 1 and 14, the suspensions did not form liquid crystals at any concentration due to aggregation of GO platelets (see S2 in ESI for further details). At pH 2, 6, and 9, a complete phase separation was observed after 4-6 days, with I phase at the top and N phase at the bottom of the capillary. As shown in Figure 1a, the ordered N phase starts to form at ~ 5 mg/ml for each of these cases, the bi-phasic (I+N) region is present at intermediate concentrations and this region broadens with increasing pH. Fully nematic (N only) phases are obtained above 16 and 18 mg/ml for pH 6 and 9 compared to 9 mg/ml at pH 2. Figures 1b-c show the retardance and azimuth images of liquid crystalline domains in the bi-phasic region formed after 1 day and after complete separation of the I and N phase, at pHs 2 and 6, respectively (for images at pH 9, please see S3 in ESI). Physically, the retardance images signify the distribution of birefringence while the azimuth images reflect the distribution of the slow axis of propagation. The ‘brush-like’ texture of the nematic phase at pH 2, seen in figure 1 (c - left) is similar to previous reports.^{6, 8, 10} But surprisingly, at pH 6 and 9 of the bi-phasic region, the nematic domains with a droplet-like near perfect spherical shape is observed for the first time as shown in figure 1(b- right).

The I-N phase transition in ideal hard platelet system is governed by the ratio of thickness (t) to the diameter (D), i.e., $\frac{t}{D}$. Thus, the concentration of onset of the I-N transition ($\Phi_{I,v}$), as well as completion ($\Phi_{N,v}$) scales with $\frac{t}{D}$. Naturally, a polydispersity of the physical dimensions will lead to broadening of the bi-phasic region.^{25,26} Unlike monodisperse systems, polydispersity of dimensions also gives rise to osmotic forces from smaller particles that causes larger particles to come closer to each other driven by depletion interactions.^{27,28} Thus, polydispersity of dimensions asides from broadening the biphasic region can also fractionate the platelets. This fractionation mechanism has been experimentally observed by van der Kooij *et al.*²⁵ and Wensink and Vroege²⁹ for polydispersity in thickness where thinner particles separate to form a nematic phase, leaving the thicker (with large $\frac{t}{D}$ ratio) ones to accumulate in the isotropic phase. A depletion mechanism was also observed by Zanchetta *et al.*³⁰ during formation of liquid crystals in mixtures of single and double- strands DNA. In our case the experiments were performed with a large polydispersity in size of the GO sheets and we vary the pH which regulates the surface charge on the platelets. The osmotic pressure, P , between two planar surfaces depend upon the surface charge density, σ , according to $\approx \frac{2\sigma^2 e^{-\kappa d}}{\epsilon\epsilon_0}$, where κ^{-1} is the Debye length, d is the distance between the surfaces and ϵ, ϵ_0 are the permittivity constants.²⁸ We hypothesized that the broadening of the biphasic region is based on fractionation of the lateral width of the GO sheets, driven by surface charges since one can assume the GO sheets to have uni-layer thickness of 0.7- 0.8 nm.^{10, 20} We argue that since the small GO sheets have higher surface charge density than larger sheets the osmotic pressure between them is increased and therefore the depletion, or the crowding of the large sheets, is enhanced. A qualitative illustration of the oxidative states of GO sheets at pH of 2, 6 and 9 and the charge density comparison between large and small sheets are shown in S4 in the ESI. As was suggested by Lerf and Klinowski¹² the basal plane on GO contains aromatic domains along with oxidized domains rich in hydroxyls and epoxy groups, while the edges are decorated with carboxylic groups, which would provide GO sheets a negative charge. In a polydisperse GO suspension, the charged edges will also yield polydispersity in

surface charge density because smaller sheets would have higher surface charge density than larger sheets and will increase with increasing pH in the range of 2-9. Without breaking this trend, we note that at pH 9 chemical changes of GO may occur, as observed by darkening of the suspension and by UV-vis spectrometry (see S5 in ESI). To test the hypothesis of fractionation, the number-average diameter (D_n) of the sheets in the isotropic and nematic phases at pH 2 and 6 was evaluated by quantitative analysis from scanning electron microscopy images (SEM, see S6 in ESI) of at least 500 sheets in each phase as shown in Figure 2. We obtained D_n of 390 nm \pm 59% and 430 nm \pm 71% for the I and N phases of pH 2, and 330 nm \pm 51% and 515 nm \pm 76% for the I and N phases of pH 6, respectively. It is clear from Figure 2 that the fractionation is enhanced at pH 6 compared to pH 2 and is consistent with our hypothesis for the broadening of the I-N phase transition.

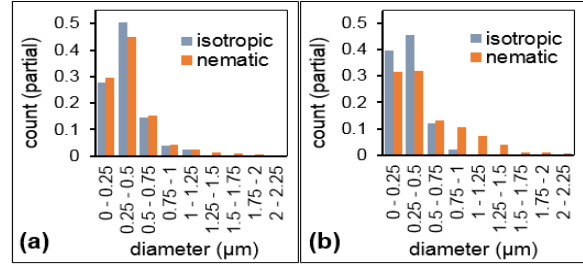


Figure 2. Size distribution of GO sheets in the isotropic and nematic phases at pH 2 (a) and pH 6 (b).

The droplets gradually merge to form a percolating ‘worm-like’ texture (images depicting the merging process are in S7 of ESI) and subsequently settles gravitationally to the bottom of the capillary as full N phase (Figure 1c-right). Liquid crystalline droplets of polymeric mesogens dispersed in continuous polymer phases have been investigated previously, but recently these droplets have demonstrated remarkable orientational sensitivity to interfacial interactions.³¹ Our spherical nematic droplets have a radial structure (Figure 3) where the director field \hat{n} is orthogonal to the droplet surface, or in other words, the GO basal plane is tangential to the surface (Figure 3d top) – the easy direction of orientation imposed by surface tension forces. Figure 3b shows the orientation of the slow axis measured with the birefringence mode and Figure 3c the orientation of high transmittance direction measured using the diattenuation mode. These two orientations are perpendicular to each other as expected from the optical characteristics of GO, having higher refractive index and higher absorbance along the basal plane. The domain has polar symmetry with a void at the center of the droplet, remarkable uniformity in diameter (20 $\mu\text{m} \pm 14\%$) obtained from 130 measurements) and an average apparent retardance of ~ 20 nm. The 2D order parameter, S , was calculated from azimuth distribution on the cross sections of the spherical domains. from ~300 pixels radially. Using $S = \langle 2\cos^2\theta - 1 \rangle$, where θ is the angle between the director and the normal of the GO sheets yields $S = 0.99$ for both pH 6 and 9. The content of the void in the center of the nematic droplet is rather intriguing. We know from birefringence (Figure 3a) and diattenuation (Figure 4a) measurements that the center of the droplet is less anisotropic than the periphery. However this could result from the presence of either isotropic GO, water or air. To resolve this issue, we looked at the average transmittance near the center, which is inversely proportional to the concentration of GO. As shown in Figure 4b, the average transmittance at the center of the droplet is similar to the transmittance at the periphery. Thus, we can conclude that the concentration of GO is similar to the isotropic or the liquid crystalline phase.

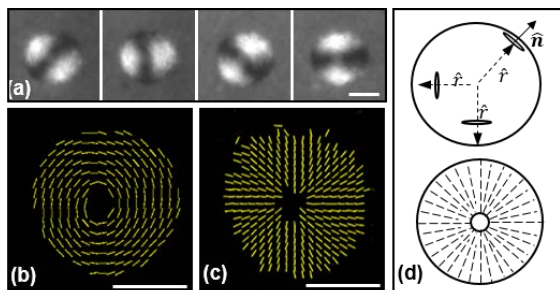


Figure 3. (a) Polarized light images of GO liquid crystalline droplet formed at pH 6 using elliptically polarized light with 45° difference in the polarization direction. (b) birefringence images displaying the orientation of the slow axis, (c) diattenuation images showing the orientation of high transmittance, (d) schematic illustration of the nematic director \hat{n} in the droplet with regard to its radius \hat{r} (top) and the radial structure of the director field (bottom). Scale bars are 10 μm .

Furthermore, we examined the role of sonication in the formation of the droplet structure to verify that it is not arising from anchoring of GO sheets around a gas bubble (air or water vapors). We found out that the nematic droplets are formed even with no sonication at all (see S9 in ESI). Because the nematic droplets are metastable, we tracked their formation path (shown in S8 in ESI). An initially isotropic suspension starts to phase separate within the first two hours (top row). Later, the nematic domains are rearranged into spherical droplets with tangential alignment of the GO sheets, followed by a second phase separation, where an isotropic phase is formed at the center of each droplet.

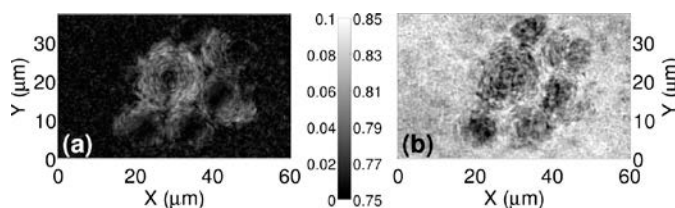


Figure 4. Diattenuation (a) and transmittance (b) of GO droplets showing that the isotropic center contains GO. The grayscale in the center shows that diattenuation varies from 0 to 0.1, whereas the transmittance with respect to clear background varies from 75% to 85% indicating presence of GO. The diameter of the isotropic center of GO droplet is at least 4 μm , whereas depth of focus of the microscope is $\sim 1.6 \mu\text{m}$.

The average diameter of the isotropic phase inside each droplet is several micron and is therefore well above the resolution limit of the microscope optics ($\lambda/(2NA) = 0.37 \mu\text{m}$). These phase separations are followed by annealing of the GO sheets inside the droplets (Figure S7, bottom row), which we presume occurs due to anchoring at the droplet interface and a directionality imposed by the surface tension.³³ However further study is required for thoroughly resolving the formation mechanism of these droplet nematic phases.

Conclusions

In this communication we report the pH dependent phase behavior of aqueous dispersions of GO for the first time. The studies reveal two major findings that are of interest. Firstly, isotropic to nematic phase transitions in these suspensions are governed by not only polydispersity of dimensions, but also polydispersity of surface charges. These two coupled phenomena enhances, particularly at $\text{pH} > \text{pK}_a$ of the

carboxylic acid groups on GO, the depletion interaction to segregate large graphene ($515 \text{ nm} \pm 76\%$) sheets into the nematic phase and smaller ($330 \text{ nm} \pm 51\%$) into the isotropic. This observation may provide a solution to size-separation of graphene sheets that has proven to be quite difficult. Secondly, we report hitherto unobserved spherical droplet nematic phases of GO suspended in isotropic GO and because larger GO sheets are more hydrophobic than smaller GO sheets¹⁶, this droplet structure bears close structural analogy to oil in water emulsions.

Notes and references

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